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Practice Packet (Unit 6: Moles & Stoichio metry

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review packet

stoichiometry and the
mole april 23rd, 2018 -

this packet is a
cumulative review of
many topics from the
year that are fair game
on the stoichiometry
test a complete answer
key is included for
students reference
topics included
nomenclature writing
formulas from names
and visa

Chemistry

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Unit 5: Moles &
Stoichiometry Practice
Packet ____1. I can
define gram-formula
mass (AKA molar
mass). Definition: gram
formula mass is the
mass of one mole of
substance ____2. Given
the chemical
symbol/formula, I can
determine how many
atoms are present.

Unit 5: Moles &
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Practice Packet

HONORS CHEMISTRY:

Unit 4 Motes

Stoichiometry Test

Review Class Pd.

$C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$
a. What is the mole ratio of oxygen to carbon dioxide?

b. How many moles of carbon dioxide are produced when 4.6 mol of oxygen react?

c. How many molecules of C_3H_8 will react with 4.6 L of oxygen?

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02 \ CBHIOH molec 2.

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Unit 4 Review

Packet Answer Key

Moles & Stoich

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Cheat Sheet

Calculating Molar Mass

1. Write out formula of compound 2.

Determine number of atoms of each element present 3. multiply

number of atoms of the element by the atomic mass of the element 4.

add up the values just

add up the values just

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determined for all of the elements present in the compound
Calculating Percent Composition 1.

Moles & Stoichiometry Cheat Sheet

Stoichiometry in chemistry is a way to account for the masses of substances going into and coming out of a chemical reaction. It involves being fluid in transforming from

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moles to grams and

grams to moles. You

will need to be

effective at unit

analysis to be able to

do this.

Unit 4-Stoichiometry

- Chemistry-2 Mr.

Nordahl

The Mole and Volume

Worksheet (DOCX 15

KB) Weekly 6

Homework (DOC 52

KB) Weekly 7

Homework (DOC 55

KB) Mole Test - Review

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Packet (DOCX 18 KB)

Mole Test - Review

Packet - Answer Key
(DOCX 27 KB)

Stoichiometry- Mole-
Mole Problems

Worksheet - Answer
Key (DOCX 16 KB)

Stoichiometry - Volume-
Volume Problems

Worksheet - Answer
Key (DOCX 18 KB)

NEED ...

**Classwork and
Homework Handouts**

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Upload Big Numbers and Chemistry At the most fundamental level, the chemist needs a unit that describes a very large quantity. One of the most well-known numbers in the study of chemistry is number of units in a mole. The number of

Unit 6: Reactions and Stoichiometry

3-4 Stoichiometry

Chapter 3 EXAMPLE

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PROBLEM: Convert Between Moles and Numbers of Atoms A sample of titanium contains 8.98×10^{25} Ti atoms. What amount of Ti, in moles, does this represent?

SOLUTION: You are asked to calculate the amount (moles) of Ti in a given sample of the metal. You are given the number of atoms of Ti in the sample. Use the equality $1 \text{ mol} = 6.022 \times 10^{23}$ atoms to

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create a ...
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Chapter 3
Stoichiometry

Science Department
Page. Mr. Davidson;
Mr. Dolgos. AP
CHEMISTRY. Unit 2:
Atomic Theory; Unit 3:
Stoichiometry; Unit 4:
Solution Stoichiometry;
Unit 5 -
Thermochemistry

Science
Department's Site /
Unit 5: MOLES &

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STOICHIOMETRY

A mole of sand would fill a cube about 32 km on a side. A mole of pennies stacked on top of each other would have about the same diameter as our galaxy, the Milky Way. A mole is a lot of things—but atoms and molecules are very tiny. One mole of carbon atoms would make a cube that is 1.74 cm on a side, small enough to carry

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**Chapter 6 -
Stoichiometry and
the Mole - CHE
105/110 ...**

Stoichiometry Packet
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somebody should go to
the books stores,
search introduction by
shop, shelf by shelf, it
is truly problematic.

022 × 1023 (1 mole)

12C atoms = 12

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Stoichiometry

grams. Assign # 1

help- Sig figs (Khan

Academy Video online

link) Notes- Unit 2 :

Worksheet- Mole

practice problems.

Mole Packet

Answers

PRACTICE PACKET Unit

5: Moles &

Stoichiometry

STUDENT* *STUDENT

- 2 - 2 GFM (Molar

Mass) Calculations -

SHOW ALL WORK!

Calculate the gram

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formula mass of the following chemicals: 1) Br₂ 2) CsOH 3 ... Use the following equation to answer the questions below:

Mr. Dolgos Regents Chemistry

Name _____ Date _____

Period _____

Stoichiometry Practice Worksheet 1) Balance the equation and then answer the following questions: _____ NaOH



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 H₂O +
 Na₂SO₄ a) How
many moles of water
will be formed if you
start with 2.54 moles
of sodium hydroxide
and you have an
excess of sulfuric acid?

Stoichiometry Practice Worksheet - Studylib

As per the
stoichiometry of the
given balanced
chemical equation, one
mole of FeCl₃ reacts

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completely with three moles of KOH. Here, we are provided with 0.025 moles of FeCl_3 , which requires $(0.025 \times 3 = 0.075)$ moles of KOH. But, only 0.050 moles of KOH are available and thus KOH is the limiting reagent of the reaction and will decide the amount of products formed in the reaction.

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